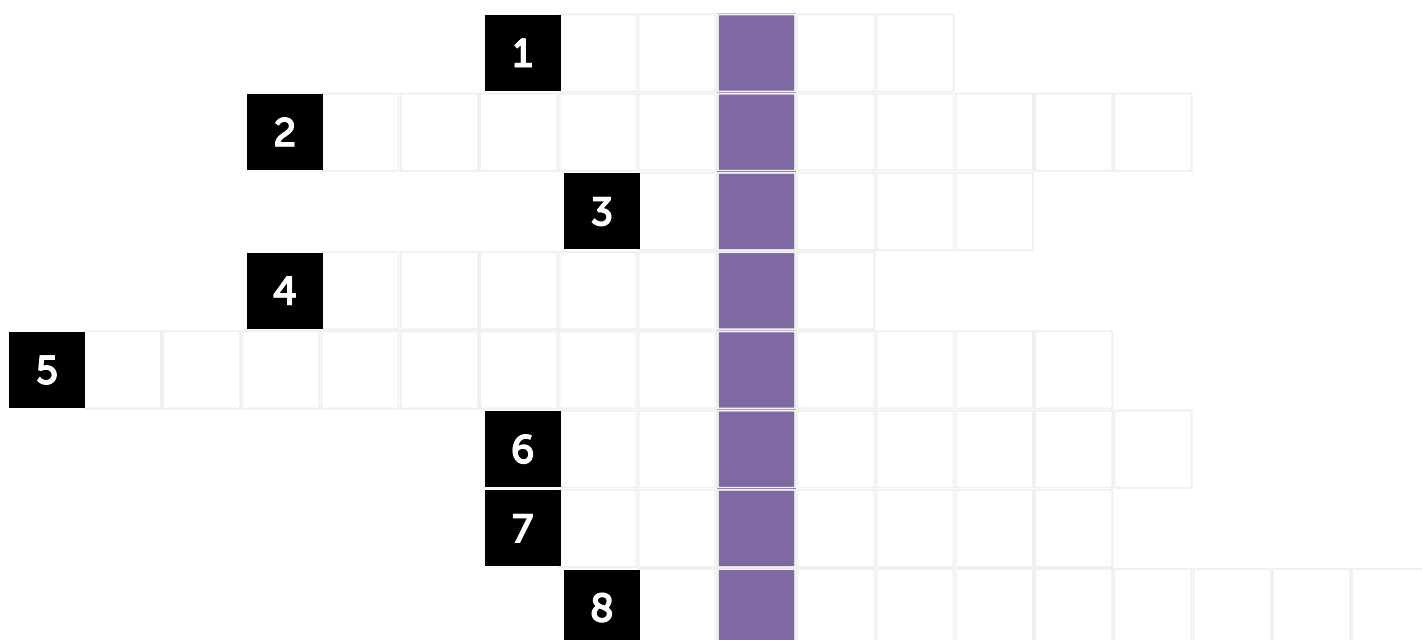


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L	E	C	A	N	D	C	J	N	C	T	A	O	O	D
F	Z	A	L	I	A	C	R	O	P	M	G	N	O	P
Q	S	Z	C	G	P	S	N	T	S	A	C	R	A	E
G	L	Z	U	K	N	O	Q	P	D	E	T	T	V	M
U	T	N	L	Q	M	L	P	R	N	I	H	J	D	U
F	U	S	A	Y	Y	D	O	T	T	W	J	K	L	L
S	L	M	T	S	Z	R	R	R	A	Y	I	M	E	O
A	E	K	I	G	A	A	A	Y	V	V	O	P	I	V
Y	Q	M	O	M	T	T	J	X	F	L	U	E	Y	G
J	D	K	N	I	I	K	G	P	E	U	C	L	Z	C
Q	N	A	O	O	M	V	C	S	F	C	B	K	E	V
Y	B	N	N	D	W	A	P	G	S	Y	M	R	Y	J
M	G	W	Y	U	T	K	Z	Q	O	J	J	H	E	W

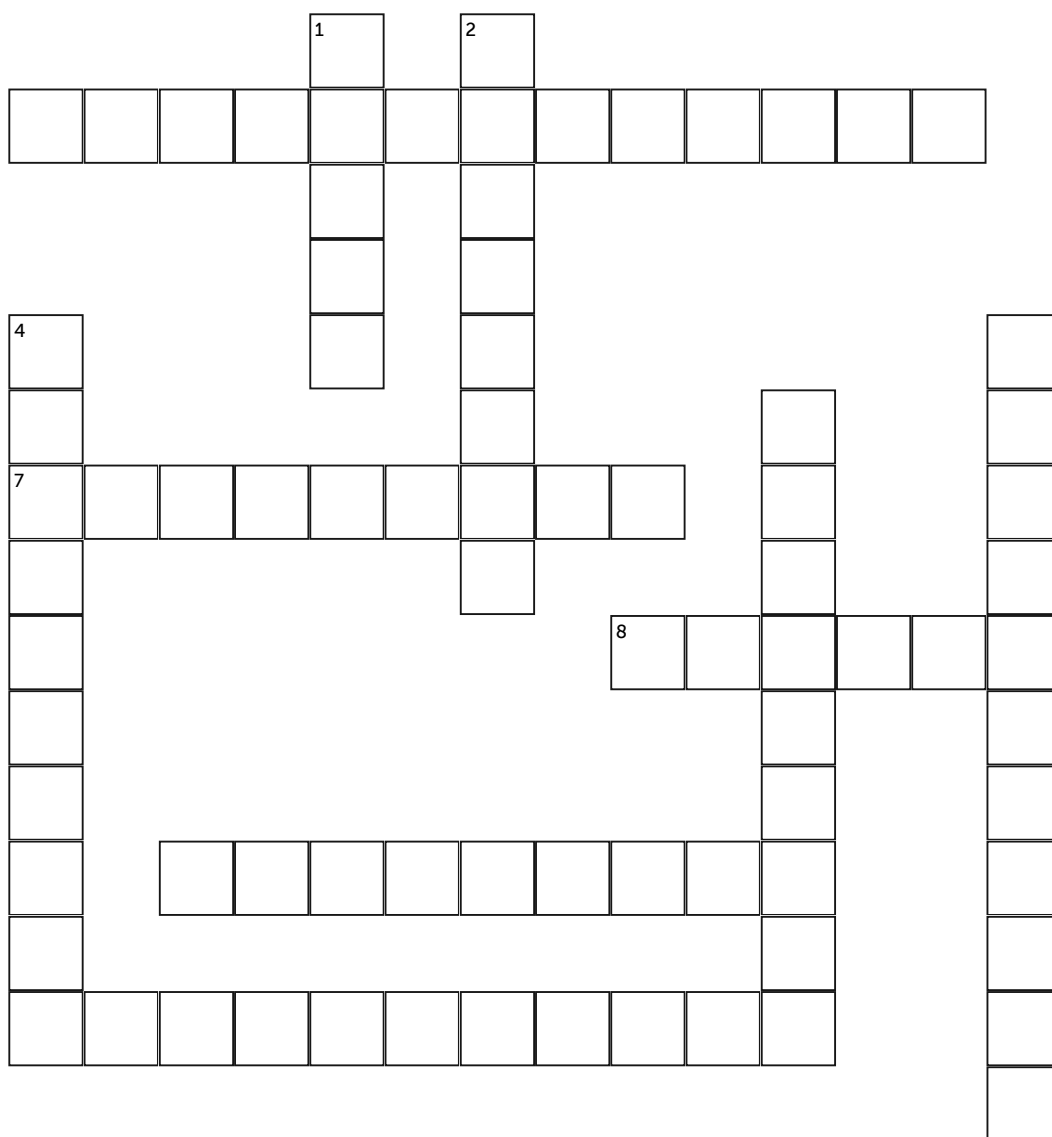
HINTS

concentration grams moles titration yield economy volume pathway
calculation avogadro



HINTS

- 1 SI unit for mass
- 2 24 dm^3 or 24000 cm^3
- 3 $\text{mass} \div \text{Mr} = ?$
- 4 We take multiple readings in experiments to calculate this
- 5 $\text{moles of solute} \div \text{volume of solvent} = ?$
- 6 Standard procedure for a titration must be carried out, such as the use of a white tile and swirling the conical flask to obtain an accurate _____
- 7 A tall piece of equipment used during titrations to accurately add a reactant
- 8 It is important that the titres you include in your calculation are within 0.10 ml



1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18	19
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ACROSS

- 3** A perfect reaction will 100% convert reactant into product, why?
- 7** Where product may be left, affecting percentage yield
- 8** _____ by 1000 to convert from cm^3 to dm^3
- 9** Avogadro's law states that equal volumes of different gases contain an equal number of _____
- 10** % yield = (actual yield \div _____ yield) $\times 100$

DOWN

- 1** Atom economy is often used to work out how much of this colour reactions are
- 2** When working with a symbol equation it is important it is this.
- 4** A reaction that only produces _____, will have a 100% atom economy.
- 5** % _____ = (RFM of desired product \div RFM of all products) $\times 100$
- 6** The indicator NOT to use during titrations