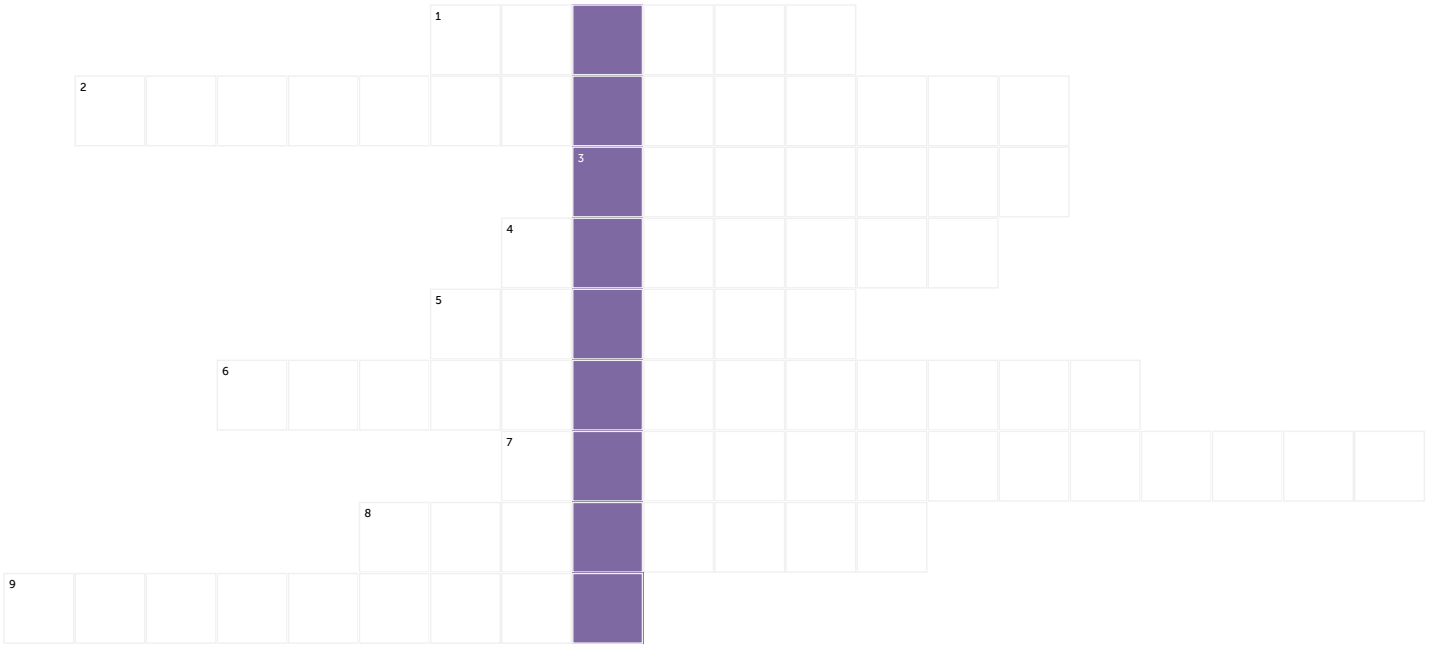
F	W	O	D	D	K	N	P	Y	E	D	D	A	T	Q
E	T	A	R	T	I	N	M	O	G	R	Z	B	G	C
L	C	W	H	A	A	G	U	O	L	F	C	Q	A	O
E	Z	E	T	F	W	L	R	L	J	Y	Z	O	B	V
C	F	D	A	V	C	O	L	A	H	C	M	V	Y	A
T	F	I	C	E	C	H	S	O	P	B	C	E	Q	L
R	U	X	A	G	I	X	D	U	T	H	G	Z	R	E
O	L	O	R	E	L	H	P	I	L	R	E	B	N	N
S	L	R	B	T	L	B	I	H	A	F	O	N	C	T
T	E	D	O	I	A	I	O	N	S	M	A	P	E	N
A	R	Y	N	H	T	E	E	X	T	H	O	T	E	I
T	E	H	A	P	E	D	S	T	X	D	M	N	E	G
I	N	I	T	A	M	Y	L	Y	Z	E	B	U	D	R
C	E	X	E	R	D	R	N	F	K	Y	B	H	R	E
H	O	L	S	G	W	I	O	N	I	C	A	E	M	Y

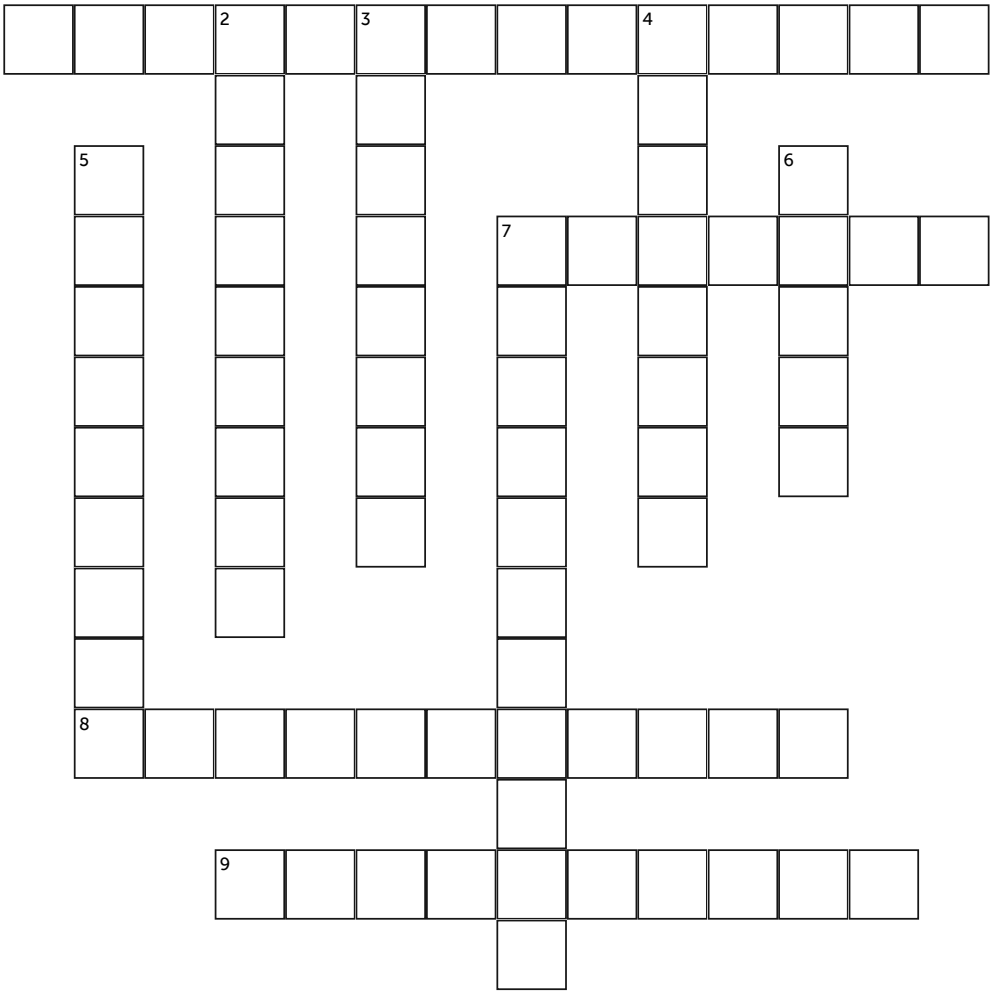
HINTS

ions hydroxide nitrate carbonate sulfate ionic covalent metallic polymer diamond graphite fullerene graphene electrostatic allotrope



HINTS

1 Electrons are _____ between atoms in covalent bonding
2 The type of weak forces that hold together covalent molecules
3 Period 2 element with 1 electron on its outershell
4 Made up of repeating units called monomers
5 A positively charged ion
6 The type of force holding metal ions together
7 The compound formed when sulfur and sodium react together
8 A form of carbon that conducts electricity
9 Can be thought of as graphene sheets rolled into a ball



1	2	3	4	5	6	7

ACROSS

1 The strength of intermolecular force between covalent molecules

7 A form of carbon that is very strong

8 Ionic compounds can conduct _____ because of their free moving ions

9 When ethene monomers join together they make this chemical

DOWN

2 Bonding happens because of these particles

3 The type of bonding in magnesium (Mg)

4 The type of bonding in water (H_2O)

5 When elements can form different structures, with different properties

6 The type of bonding in sodium chloride (NaCl)

7 The way electrons behave in a metal